

SI Chemistry
Spring 2008

Practice with Lewis Structures

Draw a Lewis Structure for each of these substances. Then determine: a) whether the bond(s) are non-polar covalent, polar covalent, or ionic, b) whether the molecule is polar or non-polar, c) the shape of the molecule and d) the intermolecular forces present. In the molecule

1. Dihydrogen sulfide
2. Carbon tetrachloride
3. Phosphorus trifluoride
4. Boron tribromide
5. Silicon dioxide
6. Hydrogen cyanide